

23/01/2026

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Memory Based Answers & Solutions

Time : 3 hrs.

for

M.M. : 300

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(Physics, Chemistry and Mathematics)

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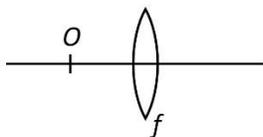
PHYSICS

SECTION - A

Multiple Choice Questions: This section contains 20 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

Choose the correct answer :

1. When an object is kept at distance 8 cm and 24 cm from a convex lens magnitude of magnification is same in both cases. Find focal length of the lens.



- (1) 32 cm (2) 8 cm
(3) 24 cm (4) 16 cm

Answer (4)

Sol. $m = m = \frac{f}{f + \mu}$

$$\frac{f}{f + 8} = \left(\frac{f}{f - 24} \right)$$

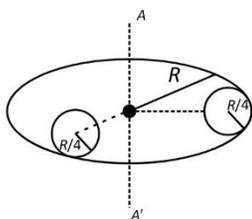
$$- f + 24 = f - 8$$

$$32 = 2f$$

$$F = 16 \text{ cm}$$

2. From a uniform disk of radius R and mass M two small disk of radius $\frac{R}{4}$ is being cut as shown in figure.

Find the moment if inertia of the system about exist AA'



- (1) $\frac{79MR^2}{128}$ (2) $\frac{79MR^2}{256}$
(3) $\frac{109MR^2}{256}$ (4) $\frac{109MR^2}{128}$

Answer (3)

Sol. $\frac{MR^2}{2} - \left(\frac{MR^2}{16 \times 2} + \frac{M9R^2}{16} \right) \times 2 = I$

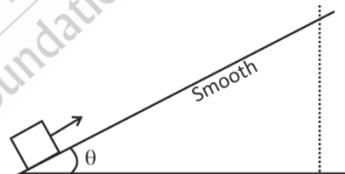
Also $m = \frac{M}{16}$

So $I = \frac{MR^2}{2} - \frac{2MR^2}{16} \left(\frac{1}{2} + 9 \right)$

$$= \frac{MR^2}{2} - \frac{2MR^2}{16 \times 16} \times \frac{19}{2}$$

$$= \frac{109}{256} MR^2$$

3. A body is projected up the smooth incline plane having angle of inclination θ with the horizontal as shown in the figure. Find the distance covered before stopping.



- (1) $\frac{u^2}{2g \cos \theta}$ (2) $\frac{u^2}{2g}$
(3) $\frac{u^2}{2g \sin \theta}$ (4) $\frac{u^2}{2g \tan \theta}$

Answer (3)

Sol. $mgh = \frac{1}{2} mu^2$

$$\frac{u^2}{2g} = h$$

$$I = \frac{h}{\sin \theta} = \frac{u^2}{2g \sin \theta}$$

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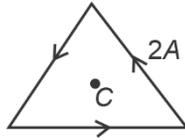
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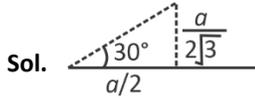
Uttar Pradesh Topper 100 Uttar Pradesh Topper 100 Telangana Topper 100

4. In equilateral triangular frame, then is current of 2A. The side of frame is $4\sqrt{3}$ cm. Magnetic field at center C is



- (1) $30\sqrt{3} \mu T$ (2) $10\sqrt{3} \mu T$
 (3) $3\sqrt{10} \mu T$ (4) $10\sqrt{10} \mu T$

Answer (1)

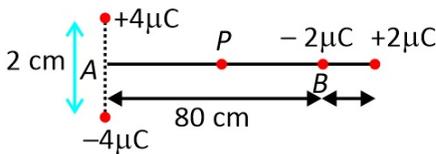


$$B = \frac{\mu_0 i \times 2 \times \text{Cu} 30^\circ}{4\pi \left\{ \frac{a}{2\sqrt{3}} \right\}}$$

$$B = \frac{\mu_0 \sqrt{3}}{2\pi \times 2 \times 10^{-2}} = 10^{-5};$$

$$B_{\text{net}} = 3\sqrt{3} \times 10^{-5}$$

5. Four charges are kept as shown in the figure. Find magnitude of electric field at point P. P is midpoint of line AB.



- (1) 180 kV/m (2) $\frac{45\sqrt{5}}{8}$ kV/m
 (3) 270 kV/m (4) $60\sqrt{3}$ kV/m

Answer (2)

Sol. $P_1 = 8 \times 10^{-8}$

$P_2 = 2 \times 10^{-8}$

$$E_1 = \frac{kP_1}{r^3}$$

$$E_2 = \frac{2kP_2}{r^3} = E_1$$

$$E_1 = \frac{9 \times 10^9 \times 8 \times 10^{-8}}{64 \times 10^{03}}, E_2 = \frac{9 \times 10^9 \times 2 \times 2 \times 10^{-8}}{64 \times 10^{-3}}$$

$$= \frac{9}{8} \times 10^4 = \frac{9}{16} \times 10^4$$

$$E = \frac{9}{16} \sqrt{5} \times 10^4$$

$$= \frac{45\sqrt{5}}{8} \times 10^3$$

6. One mole of diatomic gas is expanding isothermally from V to 2V at 27°C. If the magnitude of work done by the gas in this case is same as the work done in adiabatic process where initial temperature is 27°C and final temperature is T°C. Find the value of T. [Use $\ln 2 = 0.7$]

- (1) -37°C (2) -57°C
 (3) -35°C (4) 0°C

Answer (2)

Sol. $\Delta w_1 = NRT \ln(2) = 300R \ln(2)$

$$\Delta w_2 = N \frac{5}{2} R (T_i - T)$$

$$\Rightarrow \frac{5}{2} R (300 - T) = 300R \ln(2)$$

$$\Rightarrow 300 - T = 120 \ln(2) \quad [\ln 2 = 0.7]$$

$$\Rightarrow 300 - 84 = T = 216 \text{ K}$$

$$\Rightarrow T = -57^\circ\text{C}$$

7. Find the truth table for the given circuit.



	A	B	Y
(1)	0	0	1
	0	1	1
	1	0	0
	1	1	0
(3)	0	0	1
	0	1	0
	1	0	1
	1	1	0

	A	B	Y
(2)	0	0	0
	0	1	1
	1	0	1
	1	1	1
(4)	0	0	1
	0	1	0
	1	0	0
	1	1	1

Answer (1)

Sol. $\overline{A \cdot (A+B)}$

$$\overline{A} + \overline{(A+B)}$$

$$\overline{A} + \overline{A} \cdot \overline{B} = \overline{A}$$

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8. An air bubble is moving upward from the bottom of lake having temperature 17°C. At the top the temperature of lake (and bubble) is 27°C. Assume no significant change in the density of the lake water upto its depth of 5m. Find the ratio of volume at top to bottom of the bubble.

- (1) 1.25 (2) 1.75
(3) 1.55 (4) 1.85

Answer (3)

Sol. $P_{\text{atm}} = 10^5 \text{ N/m}^2$ $\rho_w = 10^3 \text{ kg/m}^3$

$g = 10 \text{ m/s}^2$

$$\frac{10^5}{300} \times v_T = \frac{10^5 + 5 \times 10^3 \times g \times v_B}{290}$$

$$\frac{V_T}{V_B} = 1.5 \times \frac{30}{29} = 1.55$$

9. For a medium permeability $\mu = 2\mu_0$ and dielectric constant is 3. Find the ratio of speed in vacuum to speed in medium.

- (1) $\sqrt{3}$ (2) 3
(3) $\sqrt{6}$ (4) 6

Answer (3)

Sol. $v = \frac{1}{\sqrt{\mu\epsilon}}$

$$\text{So } \frac{c}{v} = \sqrt{\frac{\mu\epsilon}{\mu_0\epsilon_0}}$$

$$\frac{c}{v} = \sqrt{6}$$

10. EMF of two cells are measured using potentiometer method. If the balance lengths are 200 cm and 150 cm respectively. If the least count is 1 cm then find % error in calculating $\frac{E_1}{E_2}$.

- (1) 1.2% (2) 1.16%
(3) 0.50% (4) 0.75%

Answer (2)

Sol. $E_1 = \left(\frac{E_0}{l_0}\right)l_1$

$$E_2 = \left(\frac{E_0}{l_0}\right)l_2$$

$$\frac{E_1}{E_2} = \frac{l_1}{l_2}$$

$$\text{So, } \left(\frac{\Delta l_1}{l_1} + \frac{\Delta l_2}{l_2}\right) \times 100 \approx 1.16 \%$$

11. A man jumps from a plane, after 2 seconds he open parachute due to which if he retarded with 3 m/s². When the man is at 10 m height from ground his speed is 5 m/s. Find height of the plane when he jumped.

- (1) 92.5 m (2) 90 m
(3) 85 m (4) 110 m

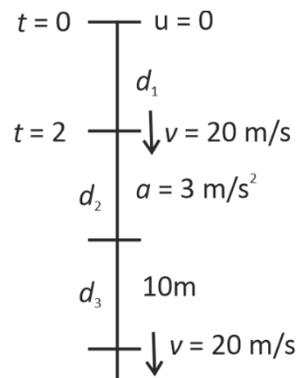
Answer (1)

Sol. $d_1 = \frac{1}{2} \times 10 \times 4 = 20 \text{ m}$

$$d_2 = \frac{400 - 25}{2 \times 3} = \frac{375}{6} = 62.5$$

$$d = 20 + 62.5 + 10$$

$$= 92.5$$



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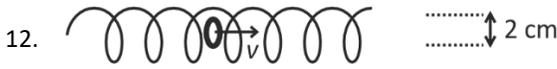


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In a long solenoid of cross-section radius of 2 cm and of 500/cm turns density. A ring moves with constant speed 10 cm/s with axis coinciding with axis of solenoid. The radius and resistance of ring is 1 cm and 10Ω . Find heat dissipated in ring while it transverses 10 cm of distance. The current in solenoid is $i = 10 \cos(100\pi t)$

- (1) $300 \mu\text{J}$ (2) $200 \mu\text{J}$
 (3) $700 \mu\text{J}$ (4) $850 \mu\text{J}$

Answer (2)

Sol. $\phi = BA$

$$\phi = \mu_0 n i A$$

$$\epsilon = -\mu_0 n A \frac{di}{dt} = \mu_0 n A i_0 \omega \sin \omega t$$

$$\int P dt = \int \frac{\mu_0^2 n^2 A^2 i_0^2 \omega^2}{R} \sin^2 \omega t dt$$

$$= \frac{\mu_0^2 n^2 A^2 i_0^2 \omega^2}{R} \times \frac{1}{2}$$

$$= \frac{(4\pi \times 10^{-7})^2 \times (50000)^2 (\pi \times 10^{-4})^2 100 \times 10^4 \times \pi^2}{10 \times 2}$$

$$= \frac{16}{2} \times 10 \times 10^{-14} \times 25 \times 10^8 \times 10 \times 10^{-8} \times 10^6 \times \frac{10}{10}$$

$$= \frac{400}{2} \times 10^{1-14+8+1-8+6}$$

$$= \frac{400}{2} \times 10^{-6} \text{ J}$$

$$= 200 \times 10^{-6}$$

13. A small spherical ball of diameter 2 mm & density 10.5 gm/cc is dropped into a large column of viscous liquid. The density of liquid is 1.5 gm/cc and coefficient of viscosity is 10 poise calculate the terminal velocity ($g = 10 \text{ m/s}^2$).

- (1) 1 cm/s (2) 2 cm/s
 (3) 3.5 cm/s (4) 1.5 cm/s

Answer (2)

Sol. $V_t = \frac{2}{9} \frac{r^2 g}{\eta} (\sigma - \rho)$

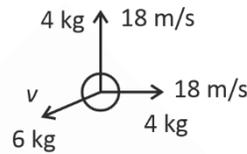
$$= 20 \times 10^{-3} \text{ m/s} = 2 \times 10^{-2} \text{ m/s}$$

14. A mass of 14 kg is exploded into three fragments of 2 : 2 : 3 and both equal fragments fly off with same speed 18 m/s in mutually perpendicular direction. Then find the speed of the third fragment immediately after the explosion.

- (1) $12\sqrt{2}$ (2) $6\sqrt{2}$
 (3) $8\sqrt{2}$ (4) $10\sqrt{2}$

Answer (1)

Sol.



$$\vec{P}_1 + \vec{P}_2 + \vec{P}_3 = 0$$

$$\vec{P}_1 + \vec{P}_2 = -\vec{P}_3$$

$$\Rightarrow 6 \times v = \sqrt{2} \times 72$$

$$V = 12\sqrt{2} \text{ m/s}$$

15. Speed of sound in air at 0°C is v . Then at what temperature ($^\circ\text{C}$) the speed of sound becomes $2v$?

- (1) 732°C (2) 1092°C
 (3) 975°C (4) 819°C

Answer (4)

Sol. $v = \sqrt{CT}$

$$\text{So } \frac{v}{v'} = \sqrt{\frac{T}{T'}}$$

$$\Rightarrow \frac{v}{2v} = \sqrt{\frac{273}{T}} = \frac{1}{2}$$

$$\Rightarrow T = 273 \times 4 = 1092 \text{ K}$$

$$\text{So, } T = 819^\circ\text{C}$$

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16. There is electric field in space given as $\vec{E} = \frac{A}{r^2} \hat{r}$ (A is constant). There are two point charges of $-2 \mu\text{C}$ and $7 \mu\text{C}$ present at $(9, 0, 0)$ and $(-9, 0, 0)$ respectively. If electric potential energy of system is zero then A in SI units is

- (1) $0.63 \times 10^{+3}$
- (2) -0.325
- (3) 1.26×10^4
- (4) 0.325

Answer (3)

Sol. $\frac{Kq}{r^2} = \frac{A}{r^2}$

$$q = \frac{A}{k}$$

$$U = \sum \frac{kq_i q_j}{r_{ij}}$$

$$U = \frac{K(-14) \times 10^{-12}}{18} + K(5) \frac{A}{k \cdot g}$$

$$= \frac{9 \times 10^9 \times 10^{-12}}{18} + K(5) \frac{A}{k \cdot g}$$

$$U = -7 \times 10^{-3} + \frac{5A \times 10^{-6}}{9}$$

$$10^6 \times \frac{63}{5} \times 10^{-3} = A$$

$$12.6 \times 10^{+3} = A$$

17. Fission of single nucleus of U-235 liberates energy of 96 MeV. Energy released by fission of 47 gm of uranium is _____.

- (1) $1.84 \times 10^{12} \text{ J}$
- (2) $3.28 \times 10^9 \text{ J}$
- (3) $1.42 \times 10^{15} \text{ J}$
- (4) $3.21 \times 10^{14} \text{ J}$

Answer (1)

Sol. Moles = $\frac{47}{235} = \frac{1}{5}$ mole

$$N = \frac{6 \times 10^{23}}{5}$$

$$E = NE = \frac{6 \times 10^{23}}{5} \times 96 \times 10^6 \times 1.6 \times 10^{-19}$$

$$= \frac{6 \times 96 \times 1.6}{5} = 10^{10}$$

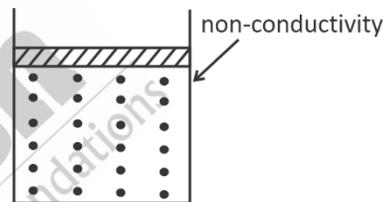
$$= 1.84 \times 10^{12} \text{ J}$$

18. 1 mole of gas is enclosed in a vertical cylinder and sealed with a massless frictionless piston as shown. On supplying 126 J of heat its temperature raised by 4°C . Find the height moved by piston due to the heat transfer. Internal energy of the gas is given by $U = 3nRT$

$$(P = 10^5 \text{ N/m}^2)$$

$$(R = 8.3 \text{ J/K})$$

$$(A_{\text{piston}} = 17 \text{ cm}^2)$$



- (1) 12.50 cm
- (2) 13.50 cm
- (3) 14.50 cm
- (4) 15.50 cm

Answer (4)

Sol. $\Delta Q = \Delta U + P\Delta V$

$$126 = 3 \times 8.3 \times 4 + 10^5 \times 16 \times 10^{-4} \times h$$

$$126 = 99.6 + 1.7 h \times 10^2$$

$$h = 15.53 \text{ cm}$$

19.

20.

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SECTION - B

Numerical Value Type Questions: This section contains 5 Numerical based questions. The answer to each question should be rounded-off to the nearest integer.

21. If r = radius of ball

ρ = density of liquid

σ = density of ball

t = time

A = constant

η = viscosity

For the expression $t = A(\rho)^a (\sigma)^b r^c (\eta)^d$

Then the value of $\frac{c-d}{a+b}$ is

Answer (3)

Sol. $T = \left(\frac{M}{L^3}\right)^{a+b} (L)^c (ML^{-1}T^{-1})^d$

Let $a + b = Z$

Then $Z + d = 0 \Rightarrow Z = -d$

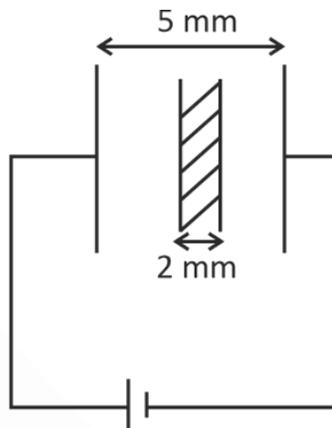
$c - 3Z - d = 0 \quad c - 2Z = 0 \Rightarrow c = 2z$

$-d = 1; Z = 1 = (a + b)$

$c = 2$

So, $\frac{c-d}{(a+b)} = \frac{2+1}{1} = 3$

22. An empty capacitor is charged with charge Q when attached with battery as shown. Now a mica sheet of 2 mm thickness is inserted and charge on capacitor $\frac{5}{4}Q$ then the dielectric constant of mica is _____.



Answer (2)

Sol. $\frac{Q_1}{Q_2} = \frac{C_1}{C_2} \Rightarrow \frac{Q}{\frac{5}{4}Q} = \frac{\frac{\epsilon_0 A}{d}}{\frac{\epsilon_0 A}{d-t+\frac{t}{K}}}$
 $\Rightarrow \frac{4}{5} = \frac{3+\frac{2}{K}}{5} \Rightarrow K=2$

23.

24.

25.

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CHEMISTRY

SECTION - A

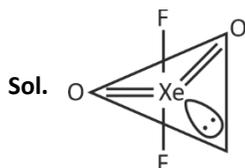
Multiple Choice Questions: This section contains 20 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

Choose the correct answer :

- For XeO_2F_2 , select the correct statement(s).
 - It shows see-saw shape.
 - Number of lone pair(s) of e^- on Xe is 1.
 - $\angle \text{FXeF} = 180^\circ$ (approx.)
 - It has tetrahedral shape.

(1) (A), (C), (D) Only (2) (A), (B) only
 (3) (A), (B), (C) only (4) (B), (C), (D) only

Answer (3)



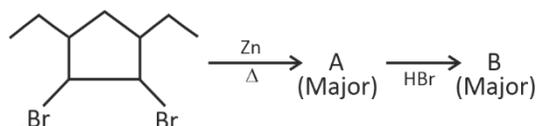
- Statement I : Size of O^{2-} is smaller than F^- .
 Statement II : Electronegativity of F is more than that of oxygen.
 In the light of above statements, choose the correct option.
 - Both Statement I and Statement II are correct
 - Both Statement I and Statement II are incorrect
 - Statement I is correct but Statement II is incorrect
 - Statement I is incorrect but Statement II is correct

Answer (4)

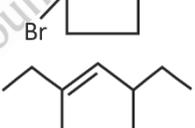
Sol. Number of proton in O^{2-} is 8 and number of proton in F^- is 9 (both have $10e^-$). So, O^{2-} is larger than F^- .
 So, Statement I is false.

Statement II is true, F is smaller than O, so F is more electronegative than O.

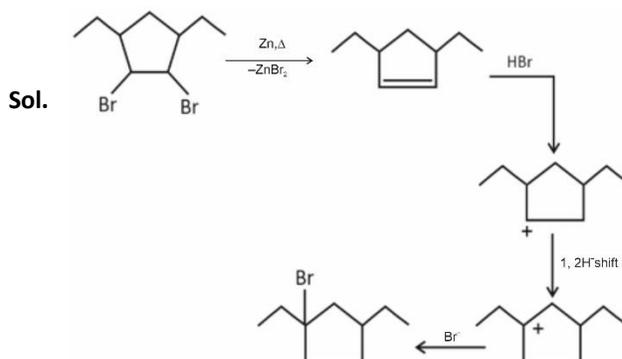
3. Consider the reaction,



Choose the correct option,

- A is 
- B is 
- B is 
- A is 

Answer (3)



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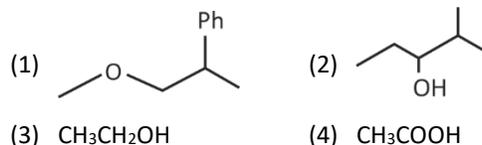
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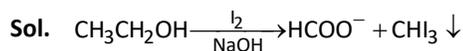
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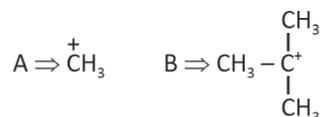
4. Which of the following molecule gives iodoform test.



Answer (3)

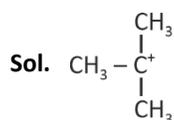


5. Consider the following intermediates.



- (1) B is more stable than A as it has 9 α hydrogen
- (2) A is more stable than B as it has 3 α hydrogen
- (3) B is more stable than A due to resonance
- (4) A is more stable due to inductive effect

Answer (1)



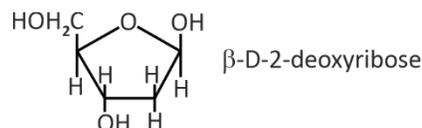
- \Rightarrow more stable as it has 9 α -H
 \Rightarrow +I effect is more than B than A

6. DNA is optically active due to the presence of

- (1) Purine nitrogenous base
- (2) Phosphate molecule
- (3) D-pentose sugar
- (4) L-pentose sugar

Answer (3)

Sol. DNA is optically active due to D-pentose sugar.



7. What is the oxidation state of chromium in the product when $\text{K}_2\text{Cr}_2\text{O}_7$ reacts with acidified KI

- (1) +6
- (2) +3
- (3) +4
- (4) +5

Answer (2)

Sol. $\text{K}_2\text{Cr}_2\text{O}_7 + 6\text{KI} + 7\text{H}_2\text{SO}_4 \rightarrow \text{Cr}_2(\text{SO}_4)_3 + 4\text{K}_2\text{SO}_4 + 3\text{I}_2 + 7\text{H}_2\text{O}$
 Cr in product have +3 oxidation state.

8. 250 cc of $x \times 10^{-3}$ M acidified $\text{K}_2\text{Cr}_2\text{O}_7$ solution titrates 750 cc of 0.6 M Mohr's salt completely. Value of x is

- (1) 200
- (2) 600
- (3) 400
- (4) 300

Answer (4)

Sol. meq of $\text{K}_2\text{Cr}_2\text{O}_7$ = meq of Mohr's salt

$$250 \times x \times 10^{-3} \times 6 = 750 \times 0.6 \times 1$$

$$\Rightarrow x = 300$$

9. Two metals with work function in ratio 1 : 2, are exposed with photons of energy 6 eV. If $\text{KE}_A : \text{KE}_B$ is 2.642 : 1, then value of ϕ_A and ϕ_B (in eV) are

- (1) 2.3, 4.6
- (2) 1.4, 2.8
- (3) 2.3, 3.6
- (4) 3.2, 6.4

Answer (1)

Sol. Let $\phi_A = Y$, then ϕ_B will be $2Y$

Let $\text{KE}_B = x$, then $\text{KE}_A = 2.642x$

$$E_i(\text{eV}) = \phi(\text{eV}) + \text{KE}(\text{eV})$$

$$\text{For metal A} \Rightarrow 6 = Y + 2.642x \quad \dots(\text{I})$$

$$\text{For metal B} \Rightarrow 6 = 2Y + x \quad \dots(\text{II})$$

$$\text{From Eq (I) and (II) } Y = 1.642x$$

Now put Y in Eq. (II)

$$6 = 2(1.642x) + x$$

$$\text{on solving } x = 1.4, Y = 1.642 \times 1.4 = 2.3$$

$$\text{So } \phi_A = Y = 2.3 \text{ eV}, \phi_B = 2Y = 4.6 \text{ eV}$$

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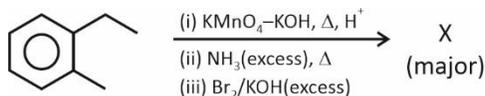
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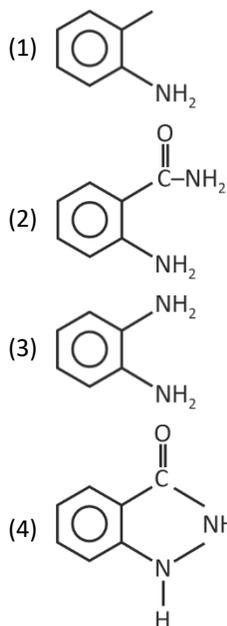
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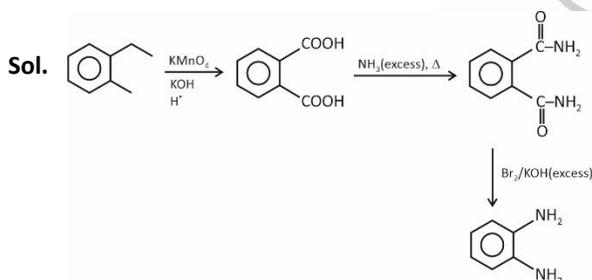
10. Consider the reaction sequence :



X is,



Answer (3)



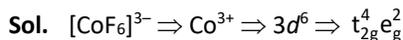
11. Consider the following statement about complexes and its hybridisation.

- A. $[\text{CoF}_6]^{3-}$; outer orbital complex, sp^3d^2
- B. $[\text{Ni}(\text{CN})_4]^{2-}$; inner orbital complex; dsp^2
- C. $[\text{Co}(\text{NH}_3)_6]^{3+}$; inner orbital complex; d^2sp^3
- D. $[\text{FeF}_6]^{3-}$; outer orbital complex; sp^3d^2

Choose the correct statement.

- (1) A, B, C and D
- (2) A, B and C only
- (3) A and B only
- (4) B and C only

Answer (1)



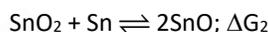
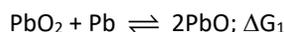
F^- is WFL $\Rightarrow sp^3d^2$, outer orbital complex.

$[\text{Ni}(\text{CN})_4]^{2-} \Rightarrow \text{Ni}^{2+}$ with $\text{CN}^- \Rightarrow \text{SFL} \Rightarrow dsp^2$, inner orbital complex

$[\text{Co}(\text{NH}_3)_6]^{3+} \Rightarrow \text{Co}^{3+} \Rightarrow \text{NH}_3$ (SFL) $\Rightarrow d^2sp^3$, inner orbital complex

$[\text{FeF}_6]^{3-} \Rightarrow \text{Fe}^{3+} \Rightarrow \text{F}^-$ (WFL) sp^3d^2 , outer orbital complex

12. Consider the following reactions



Select the correct option?

- (1) $\Delta G_1 > 0, \Delta G_2 > 0$
- (2) $\Delta G_1 < 0, \Delta G_2 > 0$
- (3) $\Delta G_1 < 0, \Delta G_2 < 0$
- (4) $\Delta G_1 > 0, \Delta G_2 < 0$

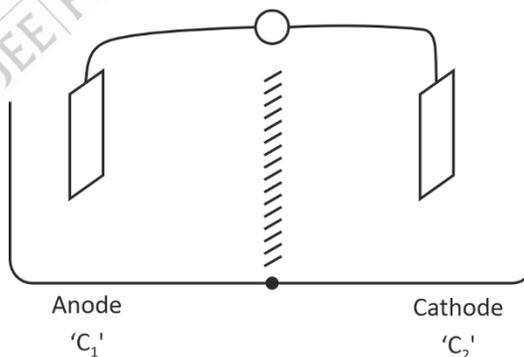
Answer (2)

Sol. Due to inert pair effect, Pb(II) is more stable than Pb(IV)

$\therefore \text{PbO} > \text{PbO}_2$ (stability)

$\text{SnO}_2 > \text{SnO}$ (stability)

13. Consider a galvanic cell, made up of two H_2 -electrodes,



Both compartments contain the same metal electrodes. If concentrations of H^+ in anode and cathode are C_1 and C_2 respectively, then $E_{\text{cell}} > 0$ when, ($p_{\text{H}_2} = 1$ atm in both compartments)

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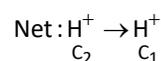
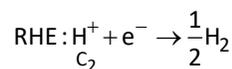
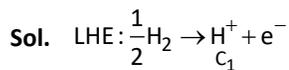


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- (1) $C_2 < C_1$
- (2) $C_2 = C_1$
- (3) $C_2 > C_1$
- (4) $C_2 = 0.5 C_1$

Answer (3)



$$E = E^\circ - \frac{0.0591}{1} \ln \frac{C_1}{C_2}$$

$$E_{\text{cell}} = 0.0591 \ln \frac{C_2}{C_1}$$

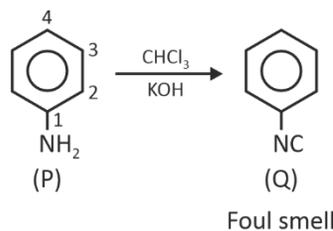
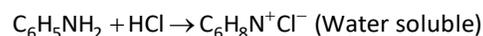
$C_2 > C_1$ for $E_{\text{cell}} = +ve$.

14. A compound P with molecular formula C_6H_7N is sparingly soluble in water. However on reaction with HCl, it becomes soluble. On reaction with $KOH + CHCl_3$, it gives foul smelling compound Q. The number of different type(s) of H atoms present in P is

- (1) 4
- (2) 5
- (3) 7
- (4) 8

Answer (1)

Sol. C_6H_7N D.U. = 4



Number of different type of H = 4

15. X & Y are elements from group 15. The difference in electronegativity values of X and phosphorus is more than difference in electronegativity of phosphorus and Y. Element X and Y respectively are

- (1) N and As
- (2) As and Sb
- (3) As and N
- (4) As and Bi

Answer (1)

Sol. EN of N = 3.0
P = 2.1
As = 2.0
Sb = 1.9
Bi = 1.9

$\Delta EN(N \text{ and } P) = 0.9$
 $\Delta EN(As \text{ and } P) = 0.1$
 $\Delta EN(Sb \text{ and } P) = 0.2$
 $\Delta EN(Bi \text{ and } P) = 0.2$

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16. Given below are two statements.

Statement I: Potassium hypiodite (KOI) can act as a reducing agent in reaction with H_2O_2 .

Statement II: When NaOCl reacts with KI in acidic medium, NaOCl acts as oxidising agent.

In the light of above statements, choose the correct option.

- (1) Both statement I and statement II are correct
- (2) Both statement I and statement II are incorrect
- (3) Statement I is correct but statement II is incorrect
- (4) Statement I is incorrect but statement II is correct

Answer (1)

Sol. $H_2O_2 + KOI \rightarrow KI + H_2O + O_2$

$NaOCl + 2KI + 2HCl \rightarrow NaCl + 2KCl + I_2 + H_2O$

- 17.
- 18.
- 19.
- 20.

SECTION - B

Numerical Value Type Questions: This section contains 5 Numerical based questions. The answer to each question should be rounded-off to the nearest integer.

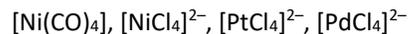
21. In estimation of chlorine by Carius method, 0.245 g organic compound gave 0.5453 g AgCl. Find percentage of chlorine in the organic compound

Answer (55)

Sol. $\% \text{ of Cl} = \frac{0.5453}{0.245} \times 35.5$
 $= \frac{143.5}{0.245} \times 100$

$= 55.06\%$

22. How many of the following complex(es) have unpaired electrons



Answer (1)

Sol. $[Ni(CO)_4] = 3d^{10} \Rightarrow sp^3$

$n = 0$

$[PtCl_4]^{2-} \Rightarrow 5d^8 \Rightarrow dsp^2$

$n = 0$

$[NiCl_4]^{2-} \Rightarrow 3d^8 \Rightarrow sp^3$

$n = 2$

$[PdCl_4]^{2-} \Rightarrow 4d^8 \Rightarrow dsp^2$

$n = 0$

23. An ideal solution is formed by mixing 3 mole of A and 1 mole of B and the vapour pressure of solution is found to be 500 mm Hg. After further addition of 1 mole A, vapour pressure of solution becomes 520 mm Hg. Find P_A^0 .

Answer (600)

Sol. $P_S = X_A P_A^0 + X_B P_B^0$

$500 = \frac{3}{4} \times P_A^0 + \frac{1}{4} \times P_B^0$

$2000 = 3P_A^0 + P_B^0 \quad \dots(i)$

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After adding 1 mole of A,

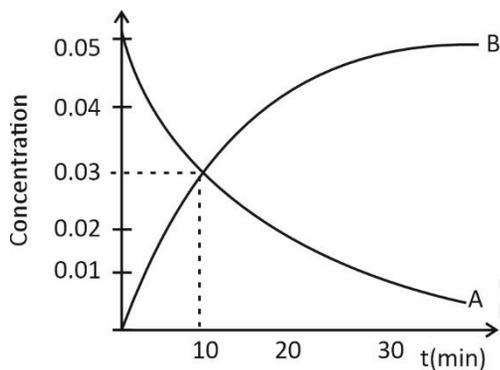
$$520 = \frac{4}{5}P_A^o + \frac{P_B^o}{5}$$

$$2600 = 4P_A^o + P_B^o \quad \dots(ii)$$

$$(ii) - (i) \Rightarrow 600 \text{ mm} = P_A^o$$

24. For a reaction $A \xrightarrow{(g)} nB_{(g)}$, a concentration vs. time

curve is,



Find 10 n

Answer (15)

Sol. For the reaction,

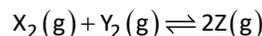
$$-\frac{dA}{dt} = +\frac{1}{n} \frac{dB}{dt}$$

$$\frac{0.05 - 0.03}{10} = \frac{1}{n} \left(\frac{0.03 - 0}{10} \right)$$

$$\frac{0.02}{0.03} = \frac{1}{n}$$

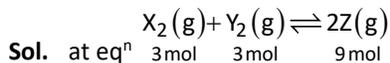
$$n = 1.5$$

25. Consider the following reaction



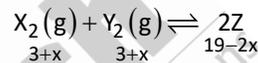
3 mol of X_2 , 3 mol of Y_2 and 9 mol of Z are present at equilibrium and the volume of container is 1L. If 10 mol of Z is added at equilibrium, calculate the number of moles of Z at new equilibrium

Answer (15)



$$K_{eq} = \frac{(9)^2}{3 \times 3} = \frac{9 \times 9}{3 \times 3} = 9$$

when 10 mol of Z is added reaction moves in backward direction



$$9 = \frac{(19-2x)^2}{(3+x)^2}$$

$$3 = \frac{19-2x}{3+x}$$

$$9 + 3x = 19 - 2x$$

$$5x = 10$$

$$x = \frac{10}{5} = 2$$

moles of Z at new equilibrium

$$= 19 - 2 \times 2$$

$$= 15$$

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MATHEMATICS

SECTION - A

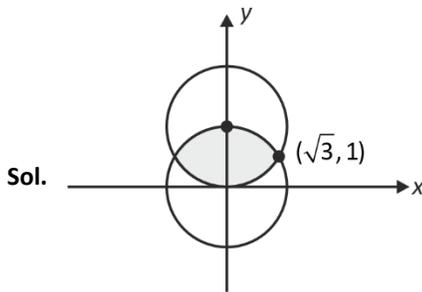
Multiple Choice Questions: This section contains 20 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

Choose the correct answer :

1. The area (in square units) between the curves $x^2 + y^2 = 4$ and $x^2 + (y - 2)^2 = 4$ is

- (1) $\frac{8\pi}{3} - 2\sqrt{3}$ (2) $\frac{8\pi}{3} + \sqrt{3}$
 (3) $\frac{4\pi}{3} - 2\sqrt{3}$ (4) $\frac{4\pi}{3} + 2\sqrt{3}$

Answer (1)



Sol.

$$A = 2 \int_0^{\sqrt{3}} (\sqrt{4-x^2} - (2-\sqrt{4-x^2})) dx$$

$$= 2 \left[\left[x\sqrt{4-x^2} + 4\sin^{-1}\left(\frac{x}{2}\right) \right]_0^{\sqrt{3}} - 2(\sqrt{3}) \right]$$

$$= 4 \left(\frac{2\pi}{3} - \frac{\sqrt{3}}{2} \right) \text{sq. units} = \frac{8\pi}{3}$$

2. Number of ways to distribute 6 identical oranges among 4 persons such that each gets at least one orange is

- (1) 8
 (2) 10
 (3) 12
 (4) 13

Answer (2)

Sol. $x_1 + x_2 + x_3 + x_4 = 6, \quad x_i \geq 1$
 $\Rightarrow x'_1 + x'_2 + x'_3 + x'_4 = 2 \quad x'_i \geq 0$
 \Rightarrow Number of ways $= {}^{2+4-1}C_{4-1}$
 $= {}^5C_3 = 10$ ways

3. The minimum value of $\cos^2 \theta + 6\sin\theta\cos\theta + 3\sin^2 \theta + 3$

- (1) -1 (2) 1
 (3) $5 + \sqrt{10}$ (4) $5 - \sqrt{10}$

Answer (4)

Sol. $\cos^2 \theta + 6\sin\theta\cos\theta + 3\sin^2 \theta + 3$
 $4 + 3\sin 2\theta + 2\sin^2 \theta$
 $4 + 3\sin 2\theta + 1 - \cos 2\theta$
 $E = 5 + 3\sin 2\theta - \cos 2\theta$
 $E_{\min} = 5 - \sqrt{10}$

4. Let $A = \{1, 2, 3, \dots, 9\}$; xRy iff $x - y$ is multiple of 3.

- S_1 : Number of elements in R is 36
 S_2 : R is equivalence relation
 (1) S_1 & S_2 both are correct
 (2) S_1 is correct but S_2 is not correct
 (3) S_2 is correct but S_1 is not correct
 (4) S_1 & S_2 both are incorrect

Answer (3)

Sol. S_1 : $(x, y) \equiv (3k_1, k_2)$ OR $(3k_1 + 1, 3k_2 + 1)$ OR $(3k_1 + 2, 3k_2 + 2)$
 $\equiv 3 \times 3 + 3 \times 3 + 3 \times 3$
 $= 27$ elements
 S_1 is incorrect.
 R is reflexive, symmetric as well as transitive relation.
 as $(x, x) \in R \quad \forall x \in A$
 $(x, y) \in R \rightarrow (y, x) \in R$
 $(x, y) \in R$ and $(y, z) \in R \Rightarrow (x, z) \in R$.

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SECTION - B

Numerical Value Type Questions: This section contains 5 Numerical based questions. The answer to each question should be rounded-off to the nearest integer.

21. If $A = \begin{bmatrix} 0 & 2 & -3 \\ -2 & 0 & -1 \\ 3 & 1 & 0 \end{bmatrix}$ and if $B(I - A) = (I + A)$, where I is

identity matrix of order 3, then trace of (BB^T) is

Answer (03.00)

Sol. $B(I - A) = (I + A)$

Since A is skew symmetric matrix

$\Rightarrow A^T = -A$ since, $B^T(I - A^T) = (I + A^T)$

$\Rightarrow B^T(I + A) = (I - A)$

$\Rightarrow BB^T(I + A) = B(I - A) = (I + A)$

$\Rightarrow BB^T(I + A)(I + A)^{-1} = (I + A)(I + A)^{-1}$

$= I_{3 \times 3}$

$\Rightarrow \text{Trace}(BB^T) = 3$

22. $S =$ No. of 4- digit numbers $abcd$ where product of digits is 20

$P =$ No. of 5-digit number $abcde$ where product of digits is 20, then

$S + P$ is equal to

Answer (74)

Sol. $\{a, b, c, d\}$

Digits : 5, 4, 1, 1 \Rightarrow No. of arrangement $\Rightarrow \frac{4!}{2!} = 12$

5, 2, 2, 1 \Rightarrow No. of arrangement $\Rightarrow \frac{4!}{2!} = 12$

$\Rightarrow S = 12 + 12 = 24$

$\{a, b, c, d, e\}$

Digits : 5, 4, 1, 1, 1 \Rightarrow No. of arrangement $\Rightarrow \frac{5!}{3!} = 20$

5, 2, 2, 1, 1 \Rightarrow No. of arrangement $\Rightarrow \frac{5!}{2!2!} = 30$

$\Rightarrow P = 20 + 30 = 50$

$S + P = 74$

23. Let S be the set defined as

$S = \left\{ x : \int_0^x t^2 \sin(t-x) dt = x^2 \text{ and } x \in [0, 1000] \right\}$, then

number of elements in S is

Answer (1)

Sol. Let $I = \int_0^x t^2 \sin(t-x) dt$

$= \int_0^x (x-t)^2 \sin((x-t)-x) dt$

$= \int_0^x (x-t)^2 \sin(-t) dt$

$= -x^2 \int_0^x (\sin t) dt + 2x \int_0^x t \sin t dt - \int_0^x t^2 \sin t dt$

$= -x^2 (-\cos t) \Big|_0^x + 2x (\sin t - t \cos t) \Big|_0^x - (2t \sin t + (2-t^2) \cos t) \Big|_0^x$

$= x^2 \cos x - x^2 + 2x(\sin x - x \cos x)$

$- 2x \sin x + (x^2 - 2) \cos x + 2$

$= x^2 \cos x - x^2 + 2x \sin x - 2x^2 \cos x - 2x \sin x$

$+ x^2 \cos x - 2 \cos x + 2$

$= -x^2 - 2 \cos x + 2$

$\Rightarrow -x^2 - 2 \cos x + 2 = x^2 \Rightarrow 2(1 - \cos x) = 2x^2$

$\Rightarrow 2 \sin^2 \frac{x}{2} = x^2 \Rightarrow x = 0$ only solution

24. The number of solutions of the equation

$\log_{(x-3)}(2x^2 - 7x + 3) = 2 \log_{(2x-1)}(x-3)$ is equal to

Answer (1)

Sol. $\log_{(x-3)}(2x^2 - 7x + 3) = 2 \log_{(2x-1)}(x-3)$

or,

$\log_{(x-3)}(x-3)(2x-1) = 2 \log_{(2x-1)}(x-3)$

or,

$1 + \log_{(x-3)}(2x-1) = 2 \log_{(2x-1)}(x-3)$

Let $\log_{(x-3)}(2x-1) = y$

$\therefore 1 + y = \frac{2}{y}$

Or, $y^2 + y - 2 = 0$

$\therefore \log_{(x-3)}(2x-1) = 1, -2$

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$$\therefore 2x - 1 = x - 3$$

$$\therefore x = -2 \text{ not acceptable.}$$

$$\text{and } 2x - 1 = (x - 3)^{-2}$$

$$\therefore (2x - 1)(x - 3)^2 = 1$$

By solving we get :

$$(2x - 5)(x^2 - 4x + 2) = 0$$

$$\therefore x = \frac{5}{2}, 2 + \sqrt{2}, 2 - \sqrt{2}$$

Only $x = 2 + \sqrt{2}$ satisfy the equations

$$\therefore \text{Number of solutions} = 1.$$

25. If a sequence $\langle a_n \rangle$ satisfy the relation $\sum_{k=1}^n a_k = \alpha n^2 + \beta n$,
and $a_{10} = 56$ and $a_6 = 2a$, then $(\alpha + \beta)$ is equal to

Answer (20)

$$\text{Sol. } \sum_{k=1}^n a_k = \alpha n^2 + \beta n$$

$$a_{10} = 56 = a + 9d$$

$$a_6 = 2a_1 \Rightarrow a + 5d = 2a$$

$$= a + 5d$$

$$\Rightarrow 14d = 56 \Rightarrow d = 4 \text{ and } a = 20$$

$$\sum_{k=1}^n a_k = \frac{n}{2}(2 \times 20 + (n-1)4)$$

$$= 20n + 2(n)(n-1)$$

$$= 2n^2 + 18n$$

$$\Rightarrow \alpha = 2, \beta = 18 \Rightarrow \alpha + \beta = 20$$



Our Problem *Solvers* shine bright in **JEE 2025**

JEE (Advanced)

ADVAY
MAYANK
AIR 36



RUJUL
GARG
AIR 41



ARUSH
ANAND
AIR 64



JEE (MAIN)

SHREYAS
LOHIYA
AIR 6
Uttar Pradesh Topper
100 Overall



KUSHAGRA
BAINGAHA
AIR 7
Uttar Pradesh Topper
100 Overall



HARSH
A GUPTA
AIR 15
Telangana Topper
100 Overall

