DATE: 21/06/2026

Test Booklet Code

60

SUSHRUT

Corporate Office: 3rd Floor, Incuspaze Campus-2, Plot No. 13,
Sector-18, Udyog Vihar, Gurugram, Haryana - 122015.

Answers & Solutions for NEET (UG)-2026 (Re-Examination)

Time : 3 hrs. 15 min.

M.M. : 720

Important Instructions:

1. The test is of **3 hours 15 minutes** duration and the Test Booklet contains **180** multiple choice questions (Four options with a single correct answer) from **Physics, Chemistry & Biology (Botany** and **Zoology).**
2. Each question carries **4 marks**. For each correct response, the candidate will get **4 marks**. For each incorrect response, **1 mark** will be deducted from the total scores. The maximum marks are **720**.
3. Use **Blue / Black Ball Point Pen only** for writing particulars on this page / marking responses on Answer Sheet.
4. Rough work is to be done in the space provided for this purpose in the Test Booklet only.
5. On completion of the test, the candidate **must handover the Answer Sheet (original & office copy) to the Invigilator** before leaving the Room / Hall. The candidates are allowed to take away this Test Booklet with them.
6. The **CODE** for this Booklet is **60**.
7. The candidates should ensure that the Answer Sheet is not folded. Do not make any stray marks on the Answer Sheet. Do not write your Roll No. anywhere else except in the specified space in the Test Booklet/Answer Sheet. Use of white fluid for correction is **NOT** permissible on the Answer Sheet.
8. Each candidate must show on-demand his/her Admit Card to the Invigilator.
9. No candidate, without special permission of the Centre Superintendent or Invigilator, would leave his/her seat.
10. Use of Electronic/Manual Calculator is prohibited.
11. The candidates are governed by all Rules and Regulations of the examination with regard to their conduct in the Examination Hall. All cases of unfair means will be dealt with as per Rules and Regulations of this examination.
12. No part of the **Test Booklet** and **Answer Sheet** shall be detached under any circumstances.
13. The candidates will write the Correct Test Booklet Code as given in the Test Booklet / Answer Sheet in the Attendance Sheet.

CHEMISTRY

46. The lanthanide ion having four unpaired electrons is

(Given : Atomic numbers of Ce = 58, Nd = 60, Tb = 65 and Ho = 67)

(1) Tb^{3+}

(2) Ho^{3+}

(3) Nd^{3+}

(4) Ce^{3+}

Answer (2)

Sol. The outermost electronic configuration is

Ho^{3+} $4f^{10}$ | ↿⇂ | ↿⇂ | ↿⇂ | 1 | 1 | 1 | 1 |

4 unpaired electrons

Nd^{3+} $4f^3$ 3 unpaired electrons

Ce^{3+} $4f^1$ 1 unpaired electron

Tb^{3+} $4f^8$ | ↿⇂ | 1 | 1 | 1 | 1 | 1 | 1 |

6 unpaired electrons

47. The standard electrode potential (E°) for the half-cell reaction $Fe^{3+} + e^- \rightarrow Fe^{2+}$ at 298 K is

(Given : $E°(Fe^{3+}/Fe) = -0.04$ V and $E°(Fe^{2+}/Fe) = -0.44$ V at 298 K)

(1) –0.48 V

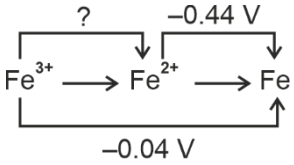
(2) +0.92 V

(3) +0.40 V

(4) +0.76 V

Answer (4)

Sol.



$\Delta G°_{Fe^{3+}/Fe} = \Delta G°_{Fe^{3+}/Fe^{2+}} + \Delta G°_{Fe^{2+}/Fe}$

$-nFE° = -nFE° + -nFE°$

$-3F(-0.04) = -1FE° + 2F(0.44)$

$3F(0.04) = -Fe° + 0.88F$

$0.12 = -E° + 0.88$

$0.12 - 0.88 = -E°$

$-0.76 = -E°$

$E° = 0.76$ V

48. Given below are two statements :

Statement-I : Heating NaCl with concentrated H_2SO_4 and MnO_2 results in oxidation of Mn.

Statement-II : Heating NaI with concentrated H_2SO_4 and MnO_2 results in reduction of Mn.

In light of the above statements, choose the **most appropriate** answer from the options given below.

(1) **Statement-I** is correct but **Statement-II** is incorrect

(2) **Statement-I** is incorrect but **Statement-II** is correct

(3) Both **Statement-I** and **Statement-II** are correct

(4) Both **Statement-I** and **Statement-II** are incorrect

Answer (2)

Sol. Statement-I : (False)

$$2NaCl + \overset{+4}{Mn}O_2 + 2H_2SO_4 \rightarrow Na_2SO_4 + \overset{+2}{Mn}SO_4 + 2H_2O + Cl_2$$

So, reduction of Mn is taking place.

Statement-II : (True)

$$2NaI + \overset{+4}{Mn}O_2 + 2H_2SO_4 \rightarrow Na_2SO_4 + \overset{+2}{Mn}SO_4 + I_2 + 2H_2O$$

So, Mn has undergone reduction.

49. The complex which has *facial* and *meridional* isomers is

(Given : py = pyridine and en = $H_2N - CH_2 - CH_2 - NH_2$)

(1) $[Co(NH_3)_4(H_2O)_2]^{3+}$

(2) $[Ni(en)_2(H_2O)_2]^{2+}$

(3) $[Cr(py)_3(Cl)_3]$

(4) $[Cr(H_2O)_6]^{3+}$

Answer (3)

Sol. The complex with type $[Ma_3b_3]$ type of complexes show fac and mer type of isomers.

$\therefore$ $[Cr(py)_3(Cl)_3]$ shows fac and mer isomers.

50. Among the following, the compound having conjugated double bonds is

(1) hepta-1,5-diene

(2) hepta-1,6-diene

(3) hepta-1,3-diene

(4) hepta-1,4-diene

Answer (3)

Sol. hepta-1,6-diene : $\overset{7}{H_2C}=\overset{6}{CH}-\overset{5}{CH_2}-\overset{4}{CH_2}-\overset{3}{CH_2}-\overset{2}{CH}=\overset{1}{CH_2}$

hepta-1,3-diene : $\overset{7}{CH_3}-\overset{6}{CH_2}-\overset{5}{CH_2}-\overset{4}{CH}=\overset{3}{CH}-\overset{2}{CH}=\overset{1}{CH_2}$

hepta-1,4-diene : $\overset{7}{CH_3}-\overset{6}{CH}-\overset{5}{CH}=\overset{4}{CH}-\overset{3}{CH_2}-\overset{2}{CH}=\overset{1}{CH_2}$

hepta-1,5-diene : $\overset{7}{CH_3}-\overset{6}{CH}=\overset{5}{CH}-\overset{4}{CH_2}-\overset{3}{CH_2}-\overset{2}{CH}=\overset{1}{CH_2}$

An organic compound (diene) with two double bonds separated by one single bond is termed as conjugated diene.

51. Match the species in **List-I** with their geometry in **List-II**.

	List-I		**List-II**
A.	PCl_5	I.	Tetrahedral
B.	BrF_5	II.	Square Planar
C.	BF_4^-	III.	Trigonal bipyramidal
D.	$[Ni(CN)_4]^{2-}$	IV.	Square pyramidal

Choose the **correct** answer from the options given below:

(1) A-III, B-I, C-II, D-IV

(2) A-III, B-II, C-I, D-IV

(3) A-IV, B-III, C-I, D-II

(4) A-III, B-IV, C-I, D-II

Answer (4)

Sol.

	Compound	**Hybridization of central atom/ion**	**Geometry**
A.	PCl_5	sp^3d	Trigonal bipyramidal
B.	BrF_5	sp^3d^2	Square pyramidal
C.	BF_4^-	sp^3	Tetrahedral
D.	$[Ni(CN)_4]^{2-}$	dsp^2	Square planar

52. The amino acid that gives a red-blood colour on treating its sodium fusion extract with sodium nitroprusside is

(1) methionine

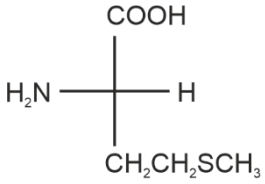
(2) serine

(3) leucine

(4) threonine

Answer (1)

Sol. In case both nitrogen and sulphur are present in an organic compound, it gives blood red colour with sodium nitroprusside.

$\therefore$ Methionine 

53. A protein undergoes reversible thermal denaturation from its initial state **N** to denatured state **D** according to $\mathbf{N} \rightleftharpoons \mathbf{D}$. At 60ºC, the concentrations of both **N** and **D** are equal at equilibrium, and the standard enthalpy change of denaturation is 666 kJ mol^{-1}. The standard entropy change (ΔS° in kJ K^{-1} mol^{-1}) of the protein upon denaturation at 60ºC is closest to

(1) 333.0

(2) 11.1

(3) 2.0

(4) 2000.0

Answer (3)

Sol. $\Delta S = \frac{\Delta H}{T}$

$= \frac{666}{333}$

= 2 kJ K^{-1} mol^{-1}

54. $2A \xrightarrow{k} B$ is a zero-order reaction, where $k = 1.0$ mol L^{-1} min^{-1}. If the initial concentration of A is 2 M, then the time taken to complete 75% of the reaction will be

(1) 1.0 min

(2) 2.0 min

(3) 1.5 min

(4) 0.75 min

Answer (4)

Sol. For zero order reaction

$$-\frac{1}{2}\frac{dA}{dt} = k$$

$$t = \frac{A_0 - A_t}{2k}$$

$$t = \frac{2 - 0.5}{2} = 0.75 \text{ min}$$

55. Given below are two statements:

Statement-I : $[Fe(ox)_3]^{3-}$ is chiral.

Statement-II : *trans* – $[Cr(H_2O)_2(ox)_2]^-$ is chiral.

(Given : oxH_2 = HOOC – COOH)

In light of the above statements, choose the **most appropriate** answer from the options given below:

(1) **Statement-I** is correct but **Statement-II** is incorrect

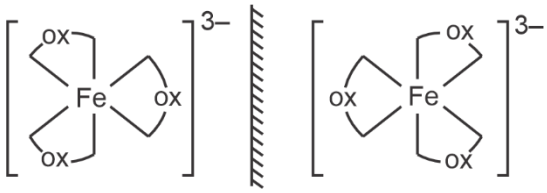
(2) **Statement-I** is incorrect but **Statement-II** is correct

(3) Both **Statement-I** and **Statement-II** are correct

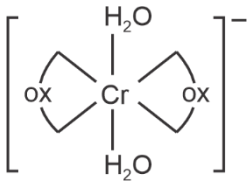
(4) Both **Statement-I** and **Statement-II** are incorrect

Answer (1)

Sol.



$[Fe(ox)_3]^{3-}$ is optically active and chiral with two non-superimposable mirror images.



Plane of symmetry (POS) is present in given species.

∴ It is NOT chiral and optically inactive.

56. Given below are two statements: One is labelled as **Assertion A** and the other is labelled as **Reason R.**

Assertion A: Generally, 3*d* transition metals have high melting points.

Reason R: Involvement of 3*d*-electrons in addition to 4*s*-electrons in the interatomic metallic bonding.

In light of the above statements, choose the ***most appropriate*** answer from the options given below:

(1) **A** is correct but **R** is not correct.

(2) **A** is not correct but **R** is correct

(3) Both **A** and **R** are correct and **R** is the correct explanation of **A**

(4) Both **A** and **R** are correct and **R** is **NOT** the correct explanation of **A**

Answer (3)

Sol. The melting points of 3d transition metals are generally high. This is attributed to the involvement of greater number of electrons from (n – 1) d in addition to the ns electrons in the interatomic metallic bonding.

57. Among the species given below, the spin-only magnetic moment is highest for

(Given: Atomic number of Ti = 22, Mn = 25, Fe = 26 and Co = 27)

(1) $[Co(NH_3)_6]^{3+}$

(2) $[Ti(H_2O)_6]^{3+}$

(3) $[Mn(CN)_6]^{3-}$

(4) $[Fe(CN)_6]^{3-}$

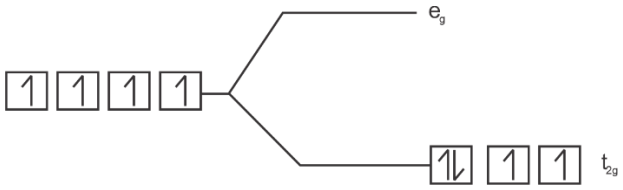
Answer (3)

Sol. In $[Ti(H_2O)_6]^{3+} \Rightarrow Ti^{3+} \Rightarrow [Ar]\ 4s^0\ 3d^1$

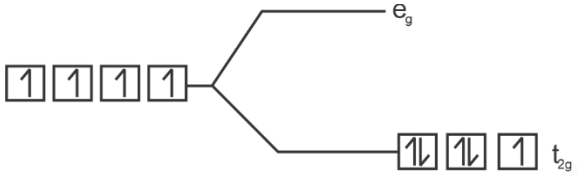
Number of unpaired electron = 1

In $[Mn(CN)_6]^{3-} \Rightarrow Mn^{3+} \Rightarrow [Ar]\ 4s^0\ 3d^4$

Number of unpaired electrons = 2



In $[Fe(CN)_6]^{3-} \Rightarrow Fe^{3+} \Rightarrow [Ar]\ 4s^0\ 3d^5$



Number of unpaired electron = 1

58. According to crystal field theory, the correct order of ligands with respect to their decreasing order of field strength is

(1) $Cl^- > H_2O > NH_3 > CO$

(2) $Cl^- > NH_3 > H_2O > CO$

(3) $CO > NH_3 > H_2O > Cl^-$

(4) $CO > H_2O > NH_3 > Cl^-$

Answer (3)

Sol. The correct decreasing order of field strength according to crystal field theory is

$CO > NH_3 > H_2O > Cl^-$

59. In potash alum, the ratio of K^+ and SO_4^{2-} ions is

(1) 2 : 3

(2) 3 : 2

(3) 1 : 2

(4) 2 : 1

Answer (3)

Sol. Potash alum

$K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24\ H_2O$

No. of $K^+ = 2$

No. of $SO_4^{2-} = 4$

Ratio $\frac{K^+}{SO_4^{2-}} = \frac{2}{4} = \frac{1}{2}$

60. Given below are two statements: One is labelled as **Assertion A** and the other is labelled as **Reason R.**

Assertion A: The first ionization enthalpy of O is lower than that of N and F.

Reason R: The loss of an electron from O leads to stable half-filled *p* orbital.

In light of the above statements, choose the ***most appropriate*** answer from the options given below:

(1) **A** is correct but **R** is not correct.

(2) **A** is not correct but **R** is correct

(3) Both **A** and **R** are correct and **R** is the correct explanation of **A**

(4) Both **A** and **R** are correct and **R** is **NOT** the correct explanation of **A**

Answer (3)

Sol. N = $1s^2\ 2s^2\ 2p^3$

F : $1s^2\ 2s^2\ 2p^5$

First ionization enthalpy (Δ_iH)

(Δ_iH) for N = 1402 kJ mol^{-1}

Δ_iH for O = 1314 kJ mol^{-1}

Δ_iH for F = 1681 kJ/mol

Due to half filled stable electronic configuration of nitrogen, it's first ionization enthalpy is higher than that of oxygen

$O(g) \longrightarrow O^+(g) + e^-$

$2s^2 2p^4 \qquad 2s^2 2p^3$

(half filled p subshell)

61. Given below are two statements:

Statement-I : Oxidation of p-nitrotoluene with acidic $KMnO_4$ gives an acid that is stronger than benzoic acid.

Statement-II : Reduction of p-nitrotoluene with Sn/HCl followed by neutralization gives an amine that is more basic than aniline.

In light of the above statements, choose the **most appropriate** answer from the options given below.

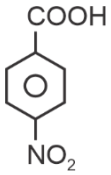
(1) **Statement-I** is correct but **Statement-II** is incorrect

(2) **Statement-I** is incorrect but **Statement-II** is correct

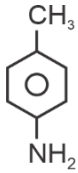
(3) Both **Statement-I** and **Statement-II** are correct

(4) Both **Statement-I** and **Statement-II** are incorrect

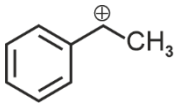
Answer (3)

Sol. • Oxidation of p-nitrotoluene with acidic $KMnO_4$ converts into p-nitrobenzoic acid.  because of −I effect and −M effect, making p-nitrobenzoic acid as stronger acid than benzoic acid.

• Reduction of p-nitrotoluene with Sn/HCl followed by neutralization reduces the ($-NO_2$) to an amino group ($-NH_2$), forming p-toluidine.

 i.e., p-methylaniline is more basic than aniline due to the +I effect and hyperconjugation.

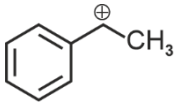
62. The following carbocation is stabilized by the interaction of the empty *p* orbital with



(1) empty σ^* and filled π orbitals
(2) empty σ^* and empty π^* orbitals
(3) filled σ and filled π orbitals
(4) empty σ and empty π^* orbitals

Answer (3)

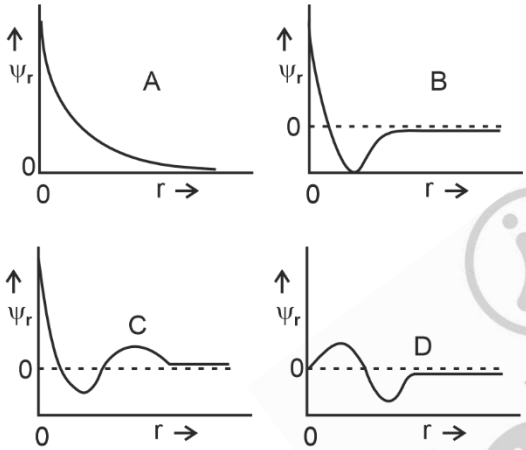
Sol.



Due to resonance it is stabilised by filled π orbital.

Due to hyperconjugation, it is stabilised by filled σ orbitals.

63. Consider the following schematic plots of orbital wavefunction (ψ_r) against distance (r) from the nucleus.

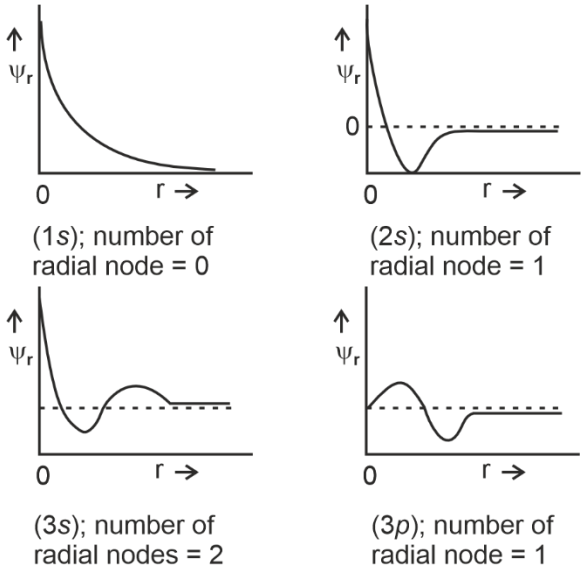


The figure representing two radial nodes in the orbital is

(1) C
(2) D
(3) A
(4) B

Answer (1)

Sol.



64. The correct order of solubility of the given salts in water at 298 K is

Salt	K_{sp} at 298 K
AgBr	5.0×10^{-13}
$Zn(OH)_2$	1.0×10^{-15}
Hg_2Cl_2	1.3×10^{-18}

(1) $Hg_2Cl_2 > AgBr > Zn(OH)_2$

(2) $Zn(OH)_2 > AgBr > Hg_2Cl_2$

(3) $Hg_2Cl_2 > Zn(OH)_2 > AgBr$

(4) $AgBr > Zn(OH)_2 > Hg_2Cl_2$

Answer (2)

Sol. $AgBr(s) \rightleftharpoons Ag^+(aq.) + Br^-(aq.)$

$K_{sp} = 5 \times 10^{-13} = S^2$

$S = \sqrt{5 \times 10^{-13}}$, $S = 7.07 \times 10^{-7}$ mol/l

$Zn(OH)_2(s) \rightleftharpoons Zn^{2+}(aq.) + 2OH^-(aq.)$

$K_{sp} = 10^{-15} = (S)(2S)^2$

$4S^3 = 10^{-15}$, $S = 0.63 \times 10^{-5} = 6.3 \times 10^{-6}$ mol/l

$Hg_2Cl_2(s) \rightleftharpoons Hg_2^{2+}(aq.) + 2Cl^-(aq.)$

$K_{sp} = 1.3 \times 10^{-18} = 4S^3$, $S = 0.69 \times 10^{-6} = 6.9 \times 10^{-7}$ mol/l

The correct order of solubility is:

$Zn(OH)_2 > AgBr > Hg_2Cl_2$

65. A 1 : 3 electrolyte in an aqueous solution is

(1) $[Co(NH_3)_6]Cl_3$

(2) $[Co(NH_3)_3(NO_2)_3]$

(3) $[CoCl_2(NH_3)_4]Cl$

(4) $[CoCl(NH_3)_5]Cl_2$

Answer (1)

Sol. The aqueous solution of $[Co(NH_3)_6]Cl_3$ is a type of 1 : 3 type electrolyte as it undergoes dissociation into $[Co(NH_3)_6]^{3+}$ and $3Cl^-$ ions.

66. The correct statement about peptides and proteins is

(1) In β-pleated sheet structures, peptide chains are held together by intermolecular hydrogen bonds

(2) In α-helices, the polypeptide chain is twisted into a left-handed screw (helix) through intramolecular hydrogen bonds

(3) Tertiary structure of proteins has two or more polypeptide subunits

(4) Only the proteins having a quaternary structure are biologically active

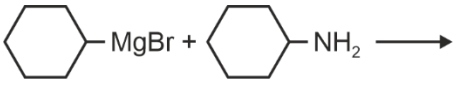
Answer (1)

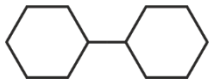
Sol. In β-pleated sheet structure, peptide chains are held together by intermolecular hydrogen bonds.

This is only correct statement.

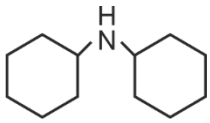
Remaining (2), (3) and (4) are incorrect statements.

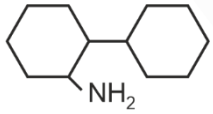
67. One of the products formed in the following reaction is



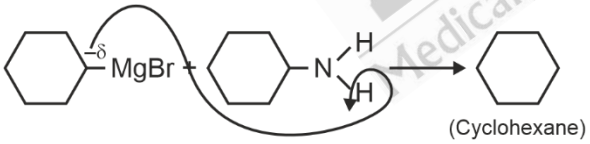
(1) 

(2) 

(3) 

(4) 

Answer (2)

Sol. 

Grignard reagents react with amines to extract proton and convert into corresponding hydrocarbon.

68. In an acidic medium, 10 mL of 0.25 M oxalic acid is titrated with $KMnO_4$ solution. If the volume of $KMnO_4$ solution required to reach end point is 10 mL, the strength of the $KMnO_4$ solution is

(1) 0.25 M

(2) 0.15 M

(3) 0.10 M

(4) 0.20 M

Answer (3)

Sol. $KMnO_4 + H_2C_2O_4 \xrightarrow{H^+} Mn^{2+} + 2CO_2 + H_2O$

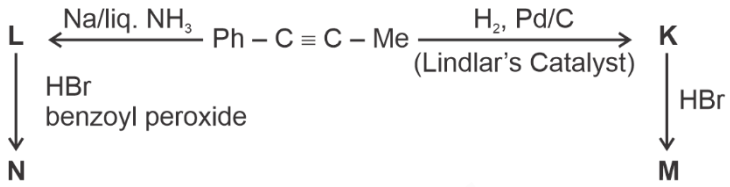
No. of equivalent of $KMnO_4$ = No. of equivalent of $H_2C_2O_4$

$10 \times 5 \times M = 10 \times 0.25 \times 2$

$M = \frac{5}{50}$

$M = \frac{1}{10} = 0.10\ M$

69. Consider the following reaction sequences and choose the correct option.



(1) **M** and **N** are geometrical isomers

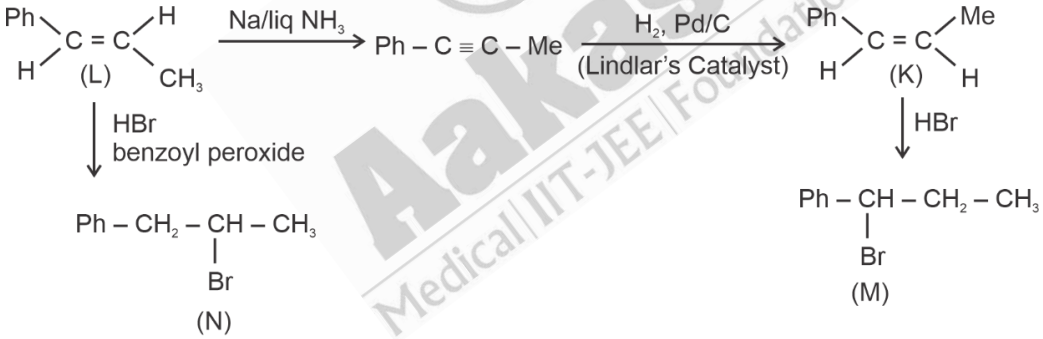
(2) **M** and **N** are stereoisomers

(3) **K** and **L** are geometrical isomers

(4) **K** and **L** are enantiomers

Answer (3)

Sol.



70. Arrange the following compounds in the increasing order of polarity

A. $CH_3CH_2OCH_2CH_3$

B. CH_3CH_2OH

C. CH_3COCH_3

D. CH_3COOH

Choose the correct answer from the options given below.

(1) C < A < B < D

(2) A < C < B < D

(3) A < B < C < D

(4) C < A < D < B

Answer (2)

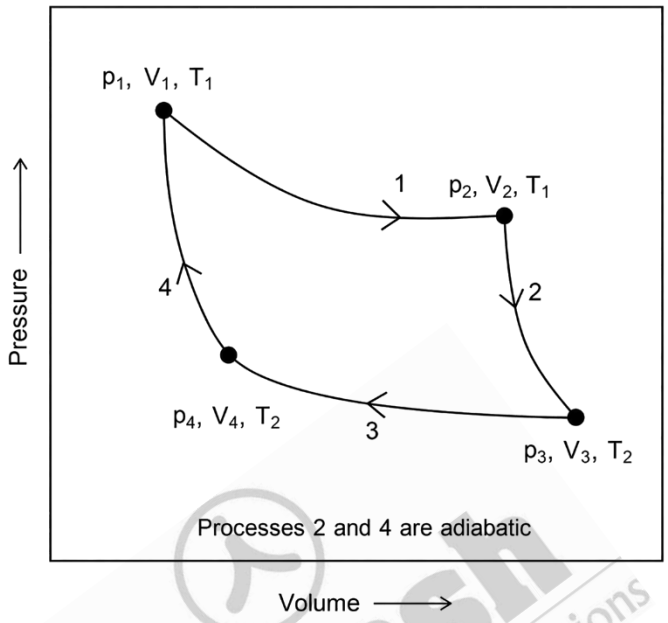
Sol. On the basis of polarity the correct order is

$C_2H_5OC_2H_5 < CH_3COCH_3 < CH_3CH_2OH < CH_3COOH$

So correct order is

A < C < B < D

71. Consider the reversible processes for 1.0 mol of an ideal gas as shown in the figure.



w_1, w_2, w_3 and w_4 represent work done (in calories) in the processes 1, 2, 3 and 4, respectively; ΔU_2 and ΔU_4 are changes in the internal energy for the processes 2 and 4, respectively.

[use R = 2 cal K^{-1} mol^{-1}]

The correct option is

(1) $w_1 + w_2 = 2T_1 \ln \frac{V_2}{V_1}$

(2) $w_1 + w_2 + w_3 + w_4 = 0$

(3) $w_1 + w_3 = -2T_1 \ln \frac{V_2}{V_1} - 2T_2 \ln \frac{V_4}{V_3}$

(4) $w_2 + w_4 = \Delta U_2 - \Delta U_4$

Answer (3)

Sol. $w_1 \rightarrow$ isothermal reversible process

$w_1 = -nRT_1 \ln \frac{V_2}{V_1}$

$w_1 = -1 \times R \times T_1 \ln \frac{V_2}{V_1}$

$w_3 = -n \times R \times T_2 \ln \frac{V_4}{V_3}$

$$w_3 = -1 \times R \times T_2 \ln \frac{V_4}{V_3}$$

$$w_1 + w_3 = -nRT_1 \ln\left(\frac{V_2}{V_1}\right) - nRT_2 \ln\left(\frac{V_4}{V_3}\right)$$

$$w_1 + w_3 = -2T_1 \ln\left(\frac{V_2}{V_1}\right) - 2T_2 \ln\left(\frac{V_4}{V_3}\right)$$

72. The green paramagnetic species formed by heating $KMnO_4$ at 513 K is

(1) MnO

(2) KO_2

(3) K_2MnO_4

(4) Mn_3O_4

Answer (3)

Sol. $2KMnO_4 \rightarrow K_2MnO_4 + MnO_2 + O_2$

K_2MnO_4 will form on heating $KMnO_4$ which has green colour.

73. The numbers 17.0145 and 21.0235 were rounded to three figures after the decimal point. The resulting numbers, respectively, are

(1) 17.014 and 21.024

(2) 17.015 and 21.024

(3) 17.014 and 21.023

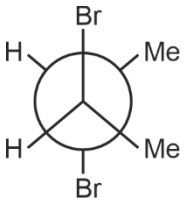
(4) 17.015 and 21.023

Answer (1)

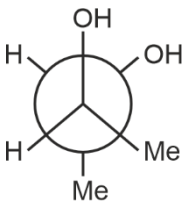
Sol. If the rightmost digit to be removed is 5, then the preceding number is not changed if it is an even number, but it is increased by one if it is an odd number.

74. Given below are two statements :

Statement I : *trans*-But-2-ene upon treatment with Br_2 in CCl_4 gives the following product.



Statement II : *cis*-But-2-ene upon treatment with alkaline $KMnO_4$ gives the following product.



In the light of the above statements, choose the ***most appropriate*** answer from the options given below.

(1) **Statement I** is correct but **Statement II** is incorrect

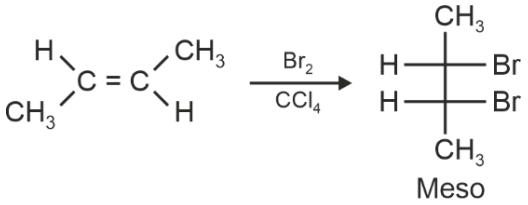
(2) **Statement I** is incorrect but **Statement II** is correct

(3) Both **Statement I** and **Statement II** are correct

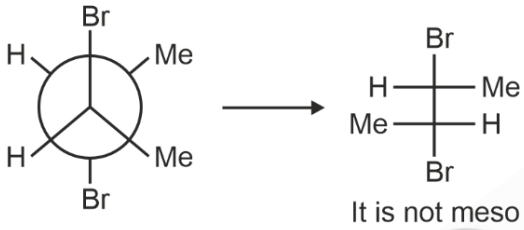
(4) Both **Statement I** and **Statement II** are incorrect

Answer (2)

Sol.

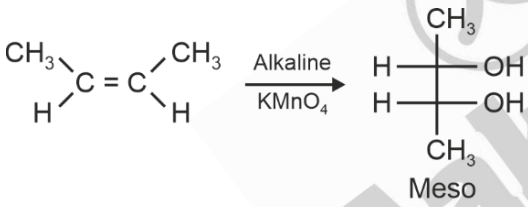


Meso

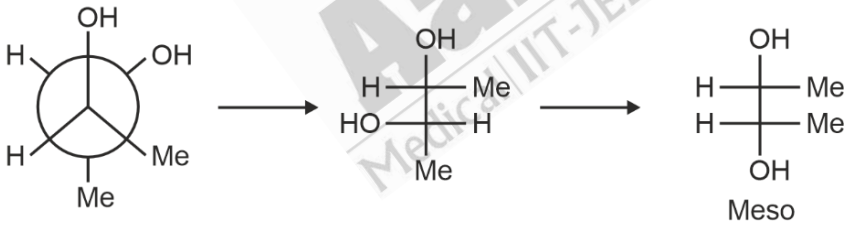


It is not meso

Therefore, statement-I is incorrect.

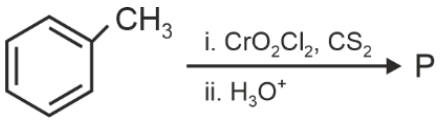


Meso



Meso

75. Consider the following reaction, and choose the correct option.



(1) On treatment with bromine water, compound **P** gives a white precipitate

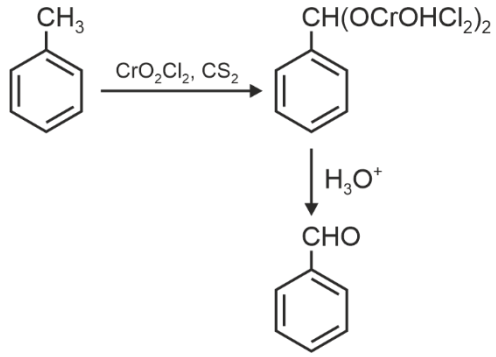
(2) Compound **P** is obtained by the hydrogenation of benzoyl chloride with Pd on $BaSO_4$

(3) On treating compound **P** with saturated $NaHCO_3$ solution, brisk effervescence is observed

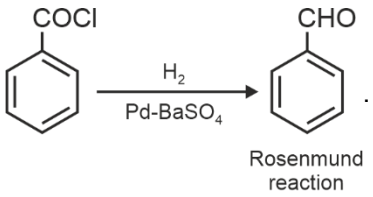
(4) Compound **P** can be prepared by treating benzene with anhydrous $AlCl_3$ and CH_3COCl

Answer (2)

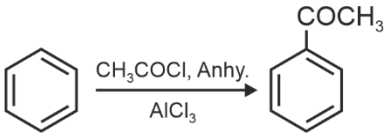
Sol.



Compound P is obtained by



Rosenmund reaction



76. Match the vitamins in **List I** with their sources in **List II**

	List I		**List II**
A.	vitamin A	I.	meat
B.	vitamin B_{12}	II.	sunflower oil
C.	vitamin E	III.	green leafy vegetables
D.	vitamin K	IV.	carrots

Choose the correct answer from the options given below.

(1) A-IV, B-II, C-I, D-III

(2) A-III, B-I, C-IV, D-II

(3) A-II, B-III, C-IV, D-I

(4) A-IV, B-I, C-II, D-III

Answer (4)

Sol.

	Vitamin	**Sources**
A.	Vitamin A	Carrot
B.	Vitamin B_{12}	Meat
C.	Vitamin E	Sunflower oil
D.	Vitamin K	Green leafy vegetables

77. The amount of carbon dioxide evolved upon complete combustion of 116 g of n-butane is

(Given: atomic mass in amu H = 1, C = 12 and O = 16)

(1) 176 g

(2) 362 g

(3) 352 g

(4) 322 g

Answer (3)

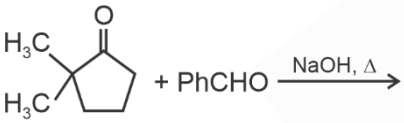
Sol. $C_4H_{10} + \frac{13}{2}O_2 \rightarrow 4CO_2 + 5H_2O$

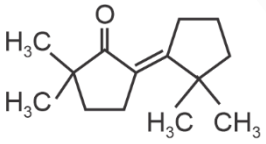
$\because$ 1 mole C_4H_{10} produces 4 mole CO_2

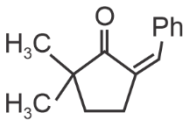
$\because$ 58 g $C_4H_{10} \equiv 4 \times 44$ g CO_2

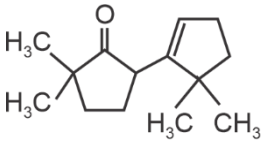
$\therefore$ 116 g $C_4H_{10} \equiv \frac{4 \times 44}{58} \times 116 = 352$ g

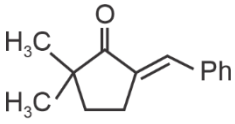
78. The compound that **CANNOT** be obtained from the aldol condensation reaction shown below, is



(1) 

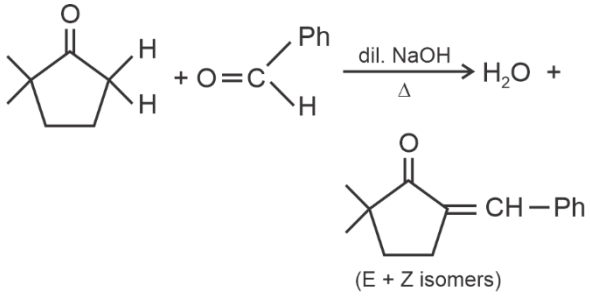
(2) 

(3) 

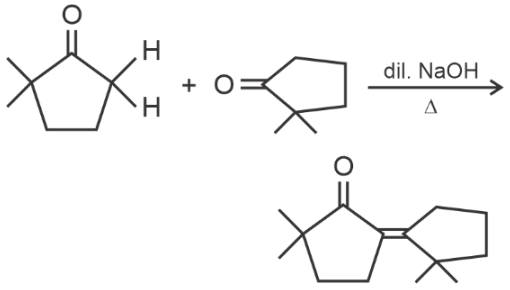
(4) 

Answer (3)

Sol.



(E + Z isomers)



α, β-unsaturated carbonyl compound

Therefore, compound 2 cannot be obtained by aldol condensation of given compounds.

79. The correct statement is

(1) Magnesium has a maximum covalency of four.

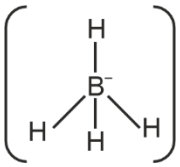
(2) Aluminium has five valence orbitals.

(3) Boron has a maximum covalency of four.

(4) Beryllium has three valence orbitals.

Answer (3)

Sol.
- Aluminium has nine valence orbitals
- Beryllium has four valence orbitals
- Magnesium has a maximum covalency of six.
- Boron has maximum covalency of four



80. The formula of tetraammineaquachloridocobalt(III) chloride is

(1) $[Co(NH_3)_4(H_2O)Cl]Cl$

(2) $[Co(NH_3)_4(H_2O)Cl]Cl_2$

(3) $[Co(NH_3)_4Cl_2] \times H_2O$

(4) $[Co(NH_3)_4]Cl_3 \times H_2O$

Answer (2)

Sol. Tetraammineaquachloridocobalt(III) chloride is $[Co(NH_3)_4(H_2O)Cl]Cl_2$.

81. **Assertion A:** For an ideal solution formed by mixing liquids **P** and **Q**, $\Delta_{mix}H = 0$ and $\Delta_{mix}V = 0$

Reason R: No interactions occur between **P** and **Q**

In the light of the above statements, choose the ***most appropriate*** answer from the options given below.

(1) **A** is correct but **R** is not correct

(2) **A** is not correct but **R** is correct

(3) Both **A** and **R** are correct and **R** is the correct explanation of **A**

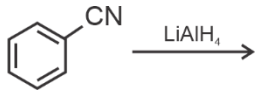
(4) Both **A** and **R** are correct but **R** is **NOT** the correct explanation of **A**

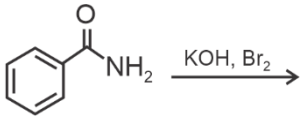
Answer (1)

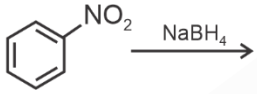
Sol. For ideal solution, $\Delta_{mix}H = 0$ and $\Delta_{mix}V = 0$

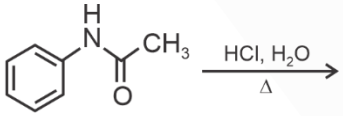
In ideal solution energy required to break interaction between P-P and Q-Q will be same as the energy released on the interaction of P-Q.

82. Identify the reactions which give aniline as the major product.

A. 

B. 

C. 

D. 

Choose the correct answer from the options given below.

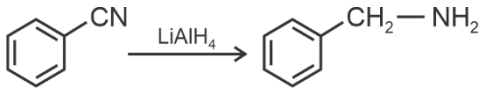
(1) A and C only

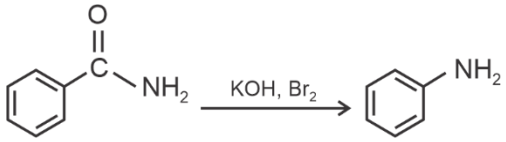
(2) C and D only

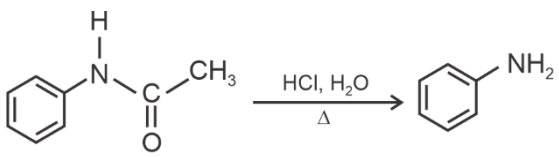
(3) A and B only

(4) B and D only

Answer (4)

Sol. 





83. The correct decreasing order of oxidation state of the underlined atom in each molecule is

(1) $\underline{Pb}O_2 > \underline{N}_2O_3 > \underline{S}O_3$

(2) $\underline{P}_4O_6 > \underline{Cl}_2O_7 > \underline{Al}H_3$

(3) $\underline{P}_4O_{10} > \underline{S}O_3 > H_2\underline{O}$

(4) $\underline{N}_2O_5 > \underline{Al}_2O_3 > H_2\underline{S}$

Answer (4)

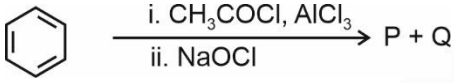
Sol. $\overset{+3}{\underline{P}}_4O_6$, $\overset{+7}{\underline{Cl}}_2O_7$, $\overset{+3}{\underline{Al}}H_3$

$\overset{+5}{\underline{P}}_4O_{10}$, $\underline{S}\overset{+6}{O}_3$, $H_2\overset{-2}{\underline{O}}$

$\overset{+5}{\underline{N}}_2O_5$, $\overset{+3}{\underline{Al}}_2O_3$, $H_2\overset{-2}{\underline{S}}$

$\overset{+4}{\underline{Pb}}_2$, $\overset{+3}{\underline{N}}_2O_3$, $\overset{+6}{\underline{S}}O_3$

84. For the following reaction sequence, choose the correct option



(1) If **P** gives a carboxylic acid on acidification, **Q** gives a poisonous gas on exposure to air and light

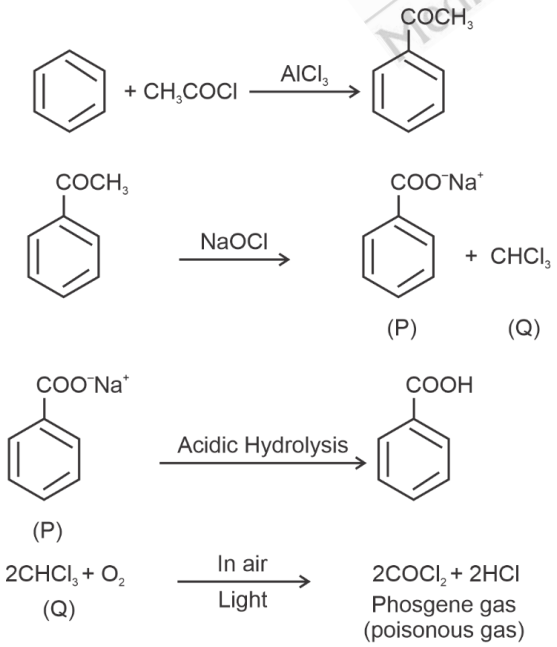
(2) Both **P** and **Q** are carbonyl compounds

(3) If **P** is the sodium salt of a carboxylic acid, **Q** is a primary alcohol

(4) **P** and **Q** are aromatic compounds

Answer (1)

Sol.



85. Two moles of an ideal gas undergo free expansion from 10 L to 100 L at 300 K. The values of ΔS_{system} and $\Delta S_{surroundings}$ are

(R is universal gas constant)

(1) $\Delta S_{system} = 0$; $\Delta S_{surroundings} = 4.606$ R

(2) $\Delta S_{system} = 4.606$ R; $\Delta S_{surroundings} = 0$

(3) $\Delta S_{system} = 0$; $\Delta S_{surroundings} = 0$

(4) $\Delta S_{system} = 4.606$ R; $\Delta S_{surroundings} = -4.606$ R

Answer (2)

Sol. During free expansion

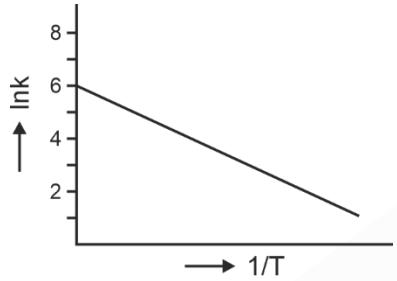
$$\Delta S_{system} = nR\ln\frac{V_2}{V_1}$$

$$= 2.303 \times 2 \times R \times \log\frac{100}{10}$$

$$= 4.606\ R$$

$\Delta S_{surroundings} = 0$ as $q = 0$

86. For an elementary chemical reaction, the Arrhenius plot is given below.



If the energy of activation is 6.64 kJ mol^{-1} and R = 8.3 J K^{-1} mol^{-1}, the temperature at which the rate constant becomes e^2 min^{-1}, is

(1) 200 K

(2) 250 K

(3) 125 K

(4) 150 K

Answer (1)

Sol. $\ln k = \ln A - \frac{E_a}{RT}$

$$\ln(e^2) = 6 - \frac{6.64 \times 10^3\ J\ mol^{-1}}{8.3\ J\ K^{-1}\ mol^{-1} \times T}$$

$$2 = 6 - \frac{6.64 \times 10^3}{8.3 \times T}$$

$$T = \frac{6.64 \times 10^3}{8.3 \times 4} = 200\ K$$

87. Consider the following statements about the solutions formed by mixing two liquids.

A. An ideal solution thus formed obeys Raoult's law throughout the composition range.

B. Mixture of chloroform and acetone shows negative deviation from Raoult's law.

C. Mixture of aniline and phenol shows positive deviation from Raoult's law.

(1) A only

(2) A and C only

(3) A and B only

(4) B and C only

Answer (3)

Sol. In the case of phenol and aniline solution, the intermolecular hydrogen bonding between phenolic proton and lone pair on nitrogen atom of aniline is stronger than the respective intermolecular hydrogen bonding between similar molecules. Therefore it shows negative deviation.

88. The highest occupied molecular orbital for Ne_2 is

(1) π^*_{2p} (2) σ^*_{2p}

(3) π_{2p} (4) σ_{2p}

Answer (2)

Sol. $Ne_2 : \sigma1s^2 < \sigma^*1s^2 < \sigma2s^2 < \sigma^*2s^2 < \sigma2p_z^2 < (\pi2p_x^2 = \pi2p_y^2) < \pi^*2p_x^2 = \pi^*2p_y^2 < \sigma^*2p_z^2$

σ^*2p_z i.e. σ^*2p is highest occupied orbital.

89. For a salt **XY**, which is a strong electrolyte, the plot of Λ_m versus $\sqrt{c}$ has a slope of –90.0 S cm^2 mol$^{-3/2}$ L$^{1/2}$ at 298 K. At 0.01 M concentration of **XY**, the value of Λ_m is 145.0 S cm^2 mol^{-1}. The limiting molar conductivity of Y^- ion ($\lambda^0_{Y^-}$, in S cm^2 mol^{-1}) at 298 K will be

(Given : $\lambda^0_{X^+}$ = 74.0 S cm^2 mol^{-1})

(1) 90.0 (2) 76.0

(3) 80.0 (4) 100.0

Answer (3)

Sol. y = mx + c

slope = –90

According to $\lambda_m = \lambda^0_m - A\sqrt{c}$

$145 = \lambda^0_m - 90 \times 0.1$

$\lambda^\circ_m = 154 \text{ S cm}^2 \text{ mol}^{-1}$

$\lambda^\circ_{m(XY)} = \lambda^\circ_{X^+} + \lambda^\circ_{Y^-}$

$154 = 74 + \lambda^\circ_{Y^-}$

$\lambda^\circ_{Y^-} = 80 \text{ S cm}^2 \text{ mol}^{-1}$

90. Among the following options, the correct trend in the electron gain enthalpy is

(1) Cl > F > Br > I

(2) I > Br > Cl > F

(3) F > Cl > Br > I

(4) Br > Cl > F > I

Answer (1)

Sol.

Group 17 Elements	Electron gain enthalpy (kJ/mol)
F	–328
Cl	–349
Br	–325
I	–295

So, the correct order of electron gain enthalpy is Cl > F > Br > I.